

Biphenyl and bimesityl tetrasulfonic acid – new linker molecules for coordination polymers

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**Dedicated to Professor Jürgen Martens in honor of his outstanding contributions to
synthetic organic chemistry**

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Abstract

Biphenyl (H₄BPTS) and bimesityl (H₄BMTS) 3,3',5,5'-tetrasulfonic acids were prepared by Ullmann type coupling reactions of respective cuprates. The four sulfo groups were installed by nucleophilic displacement at respective tetrabromo biaryl derivatives with sodium isopropyl thiolate. The resulting thioether moieties were cleaved and subsequently oxidized to furnish the title compounds. The new acids H₄BPTS and H₄BMTS were used for the preparation of their respective copper salts which allowed for the detailed structural inspection of the tetrasulfonate anions by X-ray methods.

Keywords: Biaryls, cuprates, aromatic substitution, sulfonic acids, Ullmann coupling, oxidation

Introduction

The structural chemistry of coordination polymers (CPs) is stamped by the connection of cationic knots, i.e. metal ions or metal-oxo clusters, by anionic polydentate linkers.¹⁻³ Depending on the nature of the linker the connection may occur in one, two or three dimensions. The compounds with three-dimensional linkage often show high porosity and are then called metal organic frameworks (MOFs). Due to their outstanding properties these compounds have been investigated in great detail.⁴⁻¹² In most cases CPs and MOFs have been prepared using oligocarboxylic acids as linker molecules. As an example of a tetracarboxylic acid, pyromellitic acid (**1**, Figure 1) had been utilized as a linker with *D*_{2h} symmetry, e.g. for the preparation of a material called MIL-121.¹³ An constitutionally extended example of a compound with this symmetry is 3,3',5,5'-biphenyltetracarboxylic acid (**3**), where both phenyl rings are coplanar. Its copper salt was reported as so-called MOF-505.¹⁴ Introduction of four *ortho*-methyl groups

would lead to C_{2v} symmetry of the linker, because the two aromatic rings are now perpendicular to each other. A reasonable example would be so far unknown compound **5**.

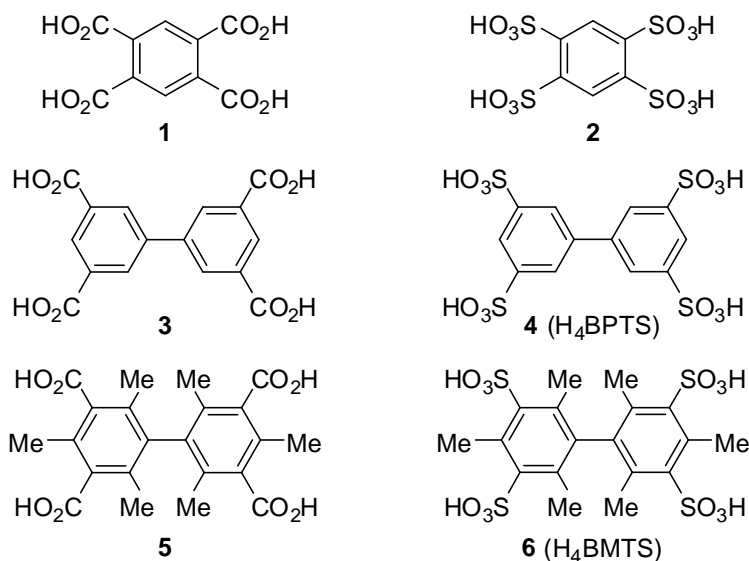
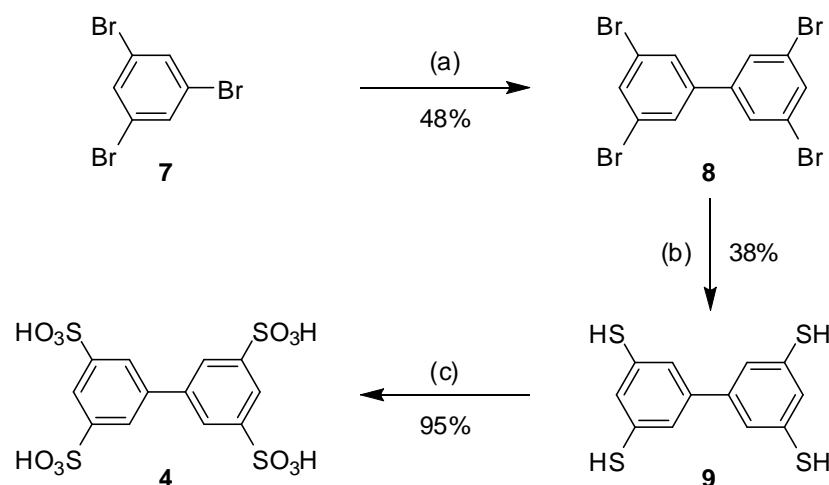


Figure 1. Aromatic tetracarboxylic and tetrasulfonic acids H_4BPTS (**4**) and H_4BMTS (**6**).

In order to extend the structural diversity of coordination polymers, we have initiated a research program on the preparation of new di- and oligosulfonic acids,^{15–19} which possess promising and advantageous properties when compared to carboxylic acids, for example, stronger acidity and enhanced thermal stability. In this course we had already reported on the use of benzene tetrasulfonic acid **2**; in the solid state, its copper(II) salt formed ladder shaped chains.²⁰ We now wish to extend this structural motif by preparation of biphenyl tetrasulfonic acid **4** (H_4BPTS) with D_{2h} symmetry and the respective tilted bimesityl congener H_4BMTS (**6**) (C_{2v}).

Results and Discussion

The synthesis of D_{2h} -symmetric tetrasulfonic acid **4** used tetrabromobiphenyl **8** as starting material, which was actually reported before in the literature;^{21,22} we however failed to reproduce the procedures used therein. Nevertheless, we could access compound **8** from tribromobenzene **7** via its Lipshutz cuprate²³ (as reported for the respective bimesityl,²⁴ see below), but we used nitrobenzene as oxidizing reagent²⁵ instead of benzoquinone. We then proceeded with the formation of tetrathiol **9** as developed before in our laboratories: nucleophilic displacement of bromine by isopropyl thioether functions was followed by reductive cleavage of isopropyl groups. Compound **9** was then oxidized by hydrogen peroxide to yield the tetrasulfonic acid **4** as its tetrahydrate ($H_4BPTS \cdot 4 H_2O$).



Reagents and conditions: (a) 1. 1.0 equiv. *n*BuLi, Et₂O, -78 °C, 10 min; 2. 0.5 equiv. CuCN, -60 °C, 1.5 h; 3. 1.5 equiv. PhNO₂, 23 °C, 16 h; (b) 10 equiv. *i*PrSNa, DMA, 100 °C, 20 h; 2. 15 equiv. Na, 100 °C, 6 h; (c) 1. 20 equiv. H₂O₂ (30% in H₂O), CHCl₃, MeOH, 23 °C, 20 h; 2. repeat step 1, but only with water as solvent.

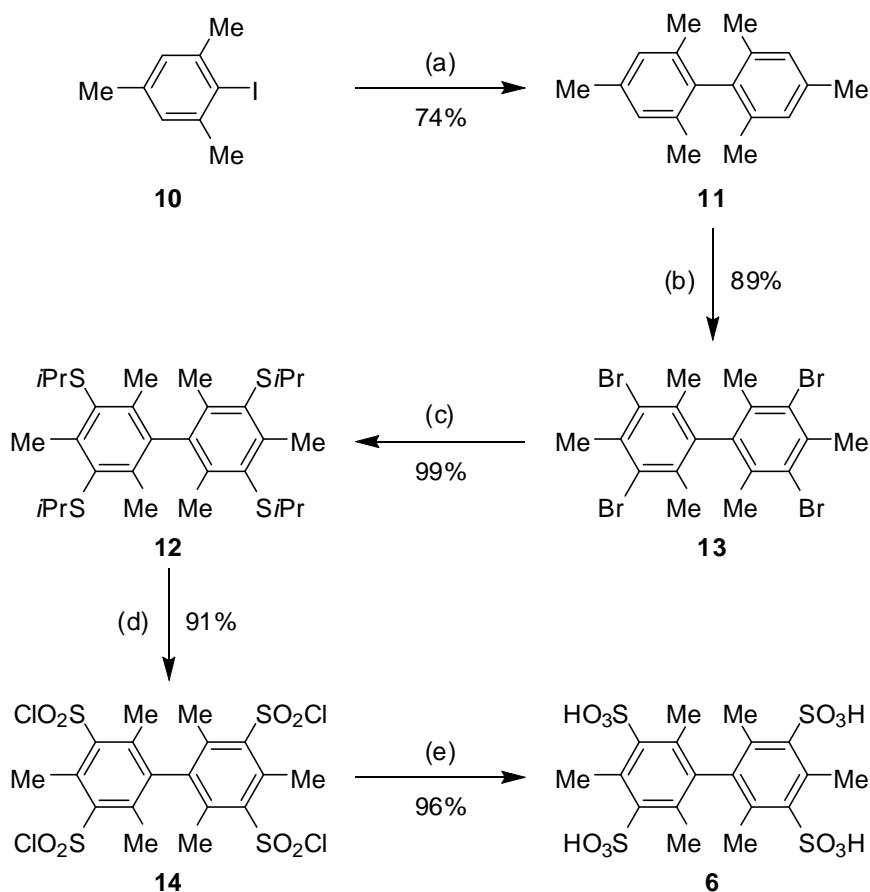
Scheme 1. Preparation of tetrasulfonic acid H₄BPTS (**4**) in three steps from tribromobenzene **7**.

For the synthesis of tetrasulfonic acid **6** we first tried the homocoupling of tribromomesitylene, but we failed. Therefore, we first prepared literature known bimesitylene^{26,27} **11** from iodomesitylene²⁸ **10** by the above approved protocol. The subsequent tetrabromination proceeded straightforward. Nucleophilic displacement of compound **13** with *i*PrSNa followed by reductive cleavage with sodium could not be performed in a sequential one flask protocol. We therefore isolated the tetrathioether **12** and degraded the isopropyl groups oxidatively with in situ formed chlorine furnishing the tetrasulfonylchloride **14**. Hydrolysis of the latter material with hot water gave the final product **6** again as the tetrahydrate (H₄BMTS · 4 H₂O).

The molecular structure of both tetrasulfonic acids H₄BPTS (**4**) and H₄BMTS (**6**) was established by X-ray crystallography of their respective copper(II) salts, which were obtained from neutralization of aqueous solutions of the respective acids with malachite followed by slow evaporation.

In the monoclinic structure of the biphenyl tetrasulfonate {[Cu(H₂O)₅]₂(BPTS)} · 2 H₂O **15** two of the [SO₃] groups are coordinated by Cu²⁺ ions both being in *anti*-conformation at the planar biphenyl backbone (Figure 2). The copper ions do not connect to further sulfonate groups and their coordination sphere is completed by five H₂O molecules. The [CuO₆] octahedron shows a typical Jahn-Teller distortion and the elongated distances are those to the sulfonate ligand (2.33 Å) and to the H₂O molecule with *trans* orientation to the [SO₃] moiety (2.41 Å). The molecular {[Cu(H₂O)₅]₂(BPTS)} complexes are connected by several hydrogen bonds,

involving also two molecules of crystal water, which are associated to each of the non-coordinating $[\text{SO}_3]$ moieties.



Reagents and conditions: (a) 1. 2.5 equiv. *t*BuLi, Et₂O, -78 °C, 30 min; 2. 0.5 equiv. CuCN, -60 °C, 2 h; 3. 1.5 equiv. PhNO₂, 23 °C, 16 h; (b) 5.5 equiv. Br₂, 0.3 equiv. Fe, CHCl₃, 23 °C, 16 h; (c) 16 equiv. *i*PrSNa, DMA, 100 °C, 20 h; (d) 16 equiv. NCS, 4.7 equiv. HCl, H₂O, MeCN, 0 °C, 20 min; (e) H₂O, THF, 100 °C, 1 d.

Scheme 2. Preparation of tetrasulfonic acid H₄BMTS (**6**) in five steps from iodomesitylene **10**.

In the tetragonal structure of the copper(II) bimesitylene tetrasulfonate $\{[\text{Cu}(\text{H}_2\text{O})_6]_2(\text{BMTS})\}$ **16** two complex $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$ cations have no coordinating contact to the tetraanion $[\text{BMTS}]^{4-}$. The octahedral complexes show a small Jahn-Teller distortion with distances of 2.028 and 2.151 Å, respectively. In Figure 3, only one of two symmetry-equivalent cations is depicted. The tetrasulfonate anion $[\text{BMTS}]^{4-}$ shows the two planes of the aromatic rings in perpendicular conformation, as is enforced by the tetra-*ortho* methyl-substitution. In the crystal structure the complex $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$ cations and the $[\text{BMTS}]^{4-}$ ions are alternately stacked in the [110] direction of the unit cell. Judged from the observed O–O distances the

formation of hydrogen bonds between the H₂O molecules of the cations and the oxygen atoms of the [SO₃] groups of the anions can be assumed.

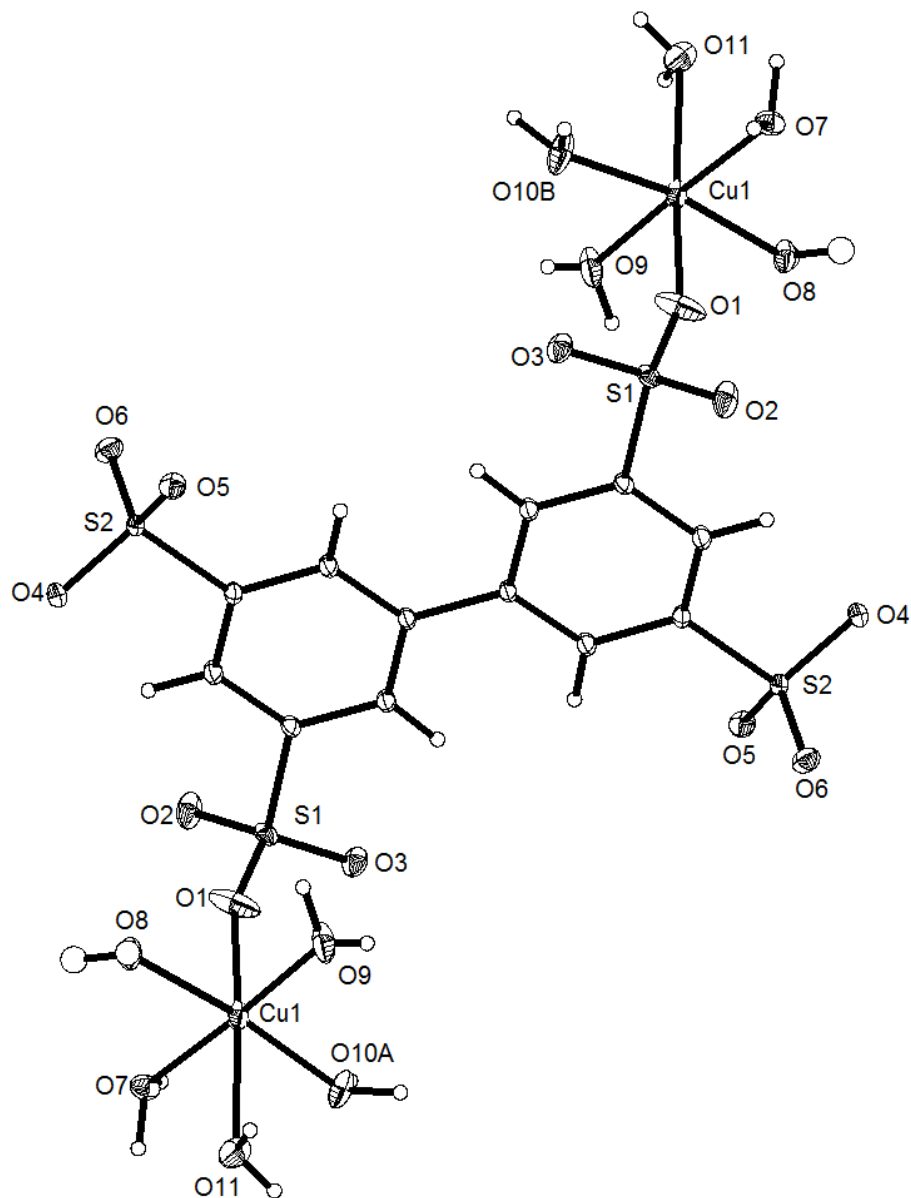


Figure 2. $\{[\text{Cu}(\text{H}_2\text{O})_5]_2(\text{BPTS})\}$ molecule in the crystal structure of the copper(II) salt **15** of the tetrasulfonic acid H₄BPTS (**4**). Thermal ellipsoids are shown at a 50% probability level.

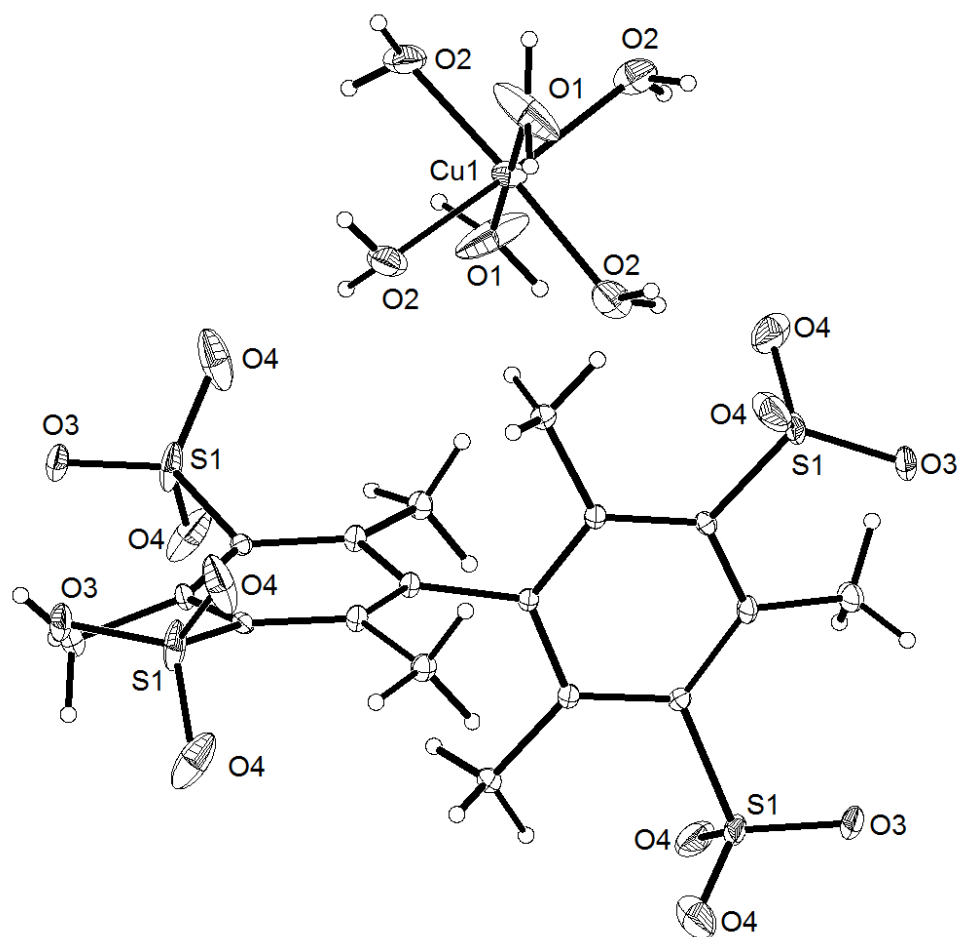


Figure 3. Tetrasulfonate anion $[\text{BMTS}]^{4-}$ and complex $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$ in the crystal structure of the copper salt $\{[\text{Cu}(\text{H}_2\text{O})_6]_2(\text{BMTS})\}$ **16** of tetrasulfonic acid H_4BMTS (**6**) in the solid state. Thermal ellipsoids are shown at a 50% probability level.

Conclusions

Two new aromatic tetrasulfonic acids with a biphenyl and bimesityl backbone, H_4BPTS (**4**) and H_4BMTS (**6**), were prepared from tetrabromobiphenyl **8** and -bimesityl **13** in two or three steps, respectively. The key transformations were Ullman-type coupling reaction to achieve the biaryl motif, nucleophilic displacement reactions with sulfur nucleophiles and oxidation to achieve the four sulfo-residues. X-ray crystallographic structure analyses of the copper salts $\{[\text{Cu}(\text{H}_2\text{O})_5]_2(\text{BPTS})\}$ **15** and $\{[\text{Cu}(\text{H}_2\text{O})_6]_2(\text{BMTS})\}$ **16** revealed the planar (D_{2h}) and tilted (C_{2v}) symmetries of the two biaryl backbones. Both tetrasulfonic acids will now be applied in our group as linker molecules for the preparation of coordination polymers and metal organic frameworks.

Experimental Section

General. Preparative column chromatography was carried out using Merck SiO₂ (35–70 μm, type 60 A) with *n*-hexane, *tert*-butyl methyl ether (MTBE), and CH₂Cl₂ as eluents. TLC was performed on Merck aluminium plates coated with SiO₂ F₂₅₄. ¹H and ¹³C NMR spectra were recorded on Bruker Avance ARX 300 and DRX 500 instruments at 23 °C in CDCl₃ or D₂O. Multiplicities of carbon signals were determined with DEPT experiments. MS and HRMS spectra were obtained with a Finnigan MAT95 (EI) and a Waters Q-TOF Premier (ESI) spectrometer. IR spectra were recorded on a Bruker Tensor 27 spectrometer equipped with a "GoldenGate" diamond ATR unit. Elemental analyses were measured with a Euro EA-CHNS instrument from HEKAtech. The X-ray single crystal investigations were performed with CCD diffractometer (κ-APEX II, Bruker, Karlsruhe, Germany). Suitable crystals were mounted onto a glass needle (Ø 0.1 mm) and placed into a stream of cold N₂ (–153 °C) inside the diffractometer. An absorption correction was applied to the data using the program SADABS-2012/1.^{29,30}

3,3',5,5'-Tetrabromobiphenyl (8). *n*-Butyllithium (2.5 mL of a 2.5 mol/l solution in *n*-hexane, 6.3 mmol) was added to a solution of 1,3,5-tribromobenzene (**7**) (2.00 g, 6.35 mmol) in Et₂O (40 mL) at –78 °C under inert atmosphere. After 10 min, CuCN (282 mg, 3.18 mmol) was added and the reaction mixture was allowed to warm up to –60 °C within 90 min. Nitrobenzene (0.98 mL, 9.53 mmol) was then added in one portion and the dark solution stirred while warming to ambient temperature (16 h). An aqueous solution of NH₃ (12.5% w/w, 80 mL) was slowly added and the resulting mixture was extracted with MTBE (3 x 50 mL). The combined organic layers were washed with water (50 mL) and brine (50 mL), dried over MgSO₄, filtered and the solvent was removed in vacuum. Further purification was accomplished by column chromatography (SiO₂, *n*-hexane, R_f 0.69). The resulting material was washed with acetone (20 mL) to give the product **8** as a colorless solid (722 mg, 1.54 mmol, 48%), mp 190 °C (ref.²² 190–191 °C). ¹H NMR (500 MHz, CDCl₃): δ 7.70 (t, *J* 1.7 Hz, 2 H), 7.59 (d, *J* 1.7 Hz, 4 H) ppm. ¹³C{¹H} NMR (126 MHz, CDCl₃): δ 141.8 (2 C), 133.9 (2 CH), 129.0 (4 CH), 123.6 (4 C) ppm. IR (ATR): 3103 (w), 3068 (m), 1579 (m), 1541 (s), 1407 (m), 1387 (m), 1098 (m), 1068 (m), 847 (s), 752 (s), 672 (m), 653 (m) cm⁻¹. MS (EI, 70 eV), *m/z* (%): 466 (95) [M⁺], 308 (25), 150 (25), 84 (100). Anal. calcd. for C₁₂H₆Br₄ (469.80): C 30.68, H 1.29; found C 30.68, H 1.30.

Biphenyl-3,3',5,5'-tetrathiol (9). Sodium 2-propanethiolate (2.76 g, 28.1 mmol) was added to a solution of compound **8** (1.32 g, 2.81 mmol) in DMA (15 mL) under inert atmosphere and the resulting suspension was stirred at 100 °C for 20 h. Sodium (969 mg, 42.1 mmol) was added at the same temperature and the reaction mixture was allowed to stir for 6 h. The conversion was monitored by GC. In case of incomplete conversion, further sodium was added. After full conversion had been achieved, water (10 mL) was slowly and carefully added and the solution was acidified by conc. hydrochloric acid until pH 1 (ca. 3 mL). More water (50 mL) was added and the resulting suspension was extracted with MTBE (3 x 50 mL). The combined organic layers were washed with hydrochloric acid (50 mL, 1 mol/l), water (50 mL) and brine (50 mL),

dried over MgSO_4 and filtered. After removal of the solvent in vacuum the residue was washed with hot methanol (100 mL) and the precipitate filtered off to yield the title compound **9** as a slightly yellow solid (466 mg, 1.65 mmol, 59%), mp 130 °C. ^1H NMR (500 MHz, CDCl_3): δ 7.18 (s, 6 H), 3.50 (s, 4 H) ppm. $^{13}\text{C}\{^1\text{H}\}$ NMR (126 MHz, CDCl_3): δ 141.4 (2 C), 132.8 (4 C), 128.7 (2 CH), 125.2 (4 CH) ppm. IR (ATR): 3047 (w), 2957 (m), 2927 (m), 2557 (m), 1573 (m), 1552 (s), 1404 (m), 1379 (m), 1127 (m), 840 (s), 795 (s), 681 (s) cm^{-1} . MS (EI, 70 eV), m/z (%): 282 (100) [M^+], 248 (8), 216 (23), 171 (10). HRMS (EI, 70 eV) calcd. for $\text{C}_{12}\text{H}_{10}\text{S}_4$: 281.9665; found 281.9668 [M^+]. Anal. calcd. for $\text{C}_{12}\text{H}_{10}\text{S}_4$ (282.45): C 51.03, H 3.57, S 45.40; found C 51.04, H 3.56, S 45.64.

Biphenyl-3,3',5,5'-tetrasulfonic acid tetrahydrate, $\text{H}_4\text{PBTS} \cdot 4 \text{H}_2\text{O}$ (4**).** Hydrogen peroxide (30% solution in H_2O , 2.17 mL, 21.2 mmol) was added to a suspension of compound **9** (300 mg, 1.06 mmol) in CHCl_3 (30 mL) and MeOH (6 mL). The reaction mixture was allowed to stir at room temperature for 20 h. The volatiles were evaporated and the residue redissolved in water (30 mL) and treated again with hydrogen peroxide (30% solution in H_2O , 2.17 mL, 21.2 mmol). After stirring for another 20 h, all volatiles were removed in vacuum to yield the product **4** as a hygroscopic colorless solid (554 mg, 1.01 mmol, 95%), decomp. 150–160 °C. ^1H NMR (500 MHz, D_2O): δ 8.28 (s, 4 H), 8.24 (s, 2 H) ppm. $^{13}\text{C}\{^1\text{H}\}$ NMR (126 MHz, D_2O): δ 144.32 (4 C), 140.6 (2 C), 127.1 (4 CH), 122.5 (2 CH) ppm. IR (ATR): 3070 (w), 1668 (s), 1110 (m), 1013 (m), 682 (m), 587 (m) cm^{-1} . MS (ESI, neg.), m/z : 473 [$\text{M} - \text{H}^+$], 236 [$\text{M} - 2 \text{H}^+$], 157 [$\text{M} - 3 \text{H}^+$]. HRMS (ESI, neg.): calcd. for $\text{C}_{12}\text{H}_8\text{O}_{12}\text{S}_4^{2-}$: 235.9455; found 235.9450 [$\text{M} - 2 \text{H}^+$]. Anal. calcd. for $\text{C}_{12}\text{H}_{10}\text{O}_{12}\text{S}_4 \cdot 4 \text{H}_2\text{O}$ (546.50): C 26.37, H 3.32, S 23.47; found C 26.37, H 3.33, S 23.45.

Bimesitylene (11**).** *tert*-Butyllithium (9.0 mL of a 1.7 mol/l solution in *n*-pentane, 15.3 mmol) was added to iodomesitylene (**10**) (1.51 g, 6.14 mmol) in Et_2O (20 mL) at -78 °C under inert atmosphere. After 30 min, CuCN (273 mg, 3.07 mmol) was added and the suspension stirred at -60 °C for 2 h. The reaction mixture was then treated with nitrobenzene (0.94 mL, 9.12 mmol) and allowed to warm up to room temperature overnight. An aqueous solution of NH_3 (12.5% w/w, 80 mL) was slowly added and the resulting residue was extracted with MTBE (3 x 50 mL). The combined organic layers were washed with water (50 mL) and brine (50 mL), dried over MgSO_4 , filtered, and the solvent was removed in vacuum. The product **11** could be obtained after column chromatography (SiO_2 , *n*-hexane, R_f 0.43) as a colorless solid (542 mg, 2.27 mmol, 74%), mp 98 °C (ref.²⁷ 99–100 °C). ^1H NMR (300 MHz, CDCl_3): δ 6.94 (s, 4 H), 2.34 (s, 6 H), 1.87 (s, 12 H) ppm. $^{13}\text{C}\{^1\text{H}\}$ NMR (75 MHz, CDCl_3): δ 137.0 (2 C), 136.0 (2 C), 135.5 (4 C), 128.2 (4 CH), 21.1 (2 CH_3), 19.8 (4 CH_3) ppm. IR (ATR): 2997 (w), 2938 (m), 2916 (m), 1470 (m), 1436 (m), 1375 (m), 1004 (m), 852 (s), 595 (m) cm^{-1} . MS (EI, 70 eV), m/z (%): 238 (99) [M^+], 223 (100), 208 (33), 193 (25). Anal. calcd. for $\text{C}_{18}\text{H}_{22}$ (238.37): C 90.70, H 9.30; found C 90.73, H 9.34.

3,3',5,5'-Tetrabromobimesitylene (13**).** A solution of bromine (0.2 mL, 3.2 mmol) in CHCl_3 (2.9 mL) was added to a suspension of bimesitylene (**11**) (140 mg, 587 μmol) and iron (9.8 mg, 176 μmol) in CHCl_3 (3.0 mL) at ambient temperature. The reaction mixture was allowed to stir for 16 h. Water (20 mL) was then added and the aqueous layer was extracted with CH_2Cl_2 (3 x

20 mL). The combined organic layers were washed with a saturated solution of $\text{Na}_2\text{S}_2\text{O}_3$ (20 mL) and water (20 mL), dried over MgSO_4 , filtered and the solvent was removed in vacuum. The crude product could be purified by washing with hot acetone (10 mL) to yield the title compound **13** as a colorless solid (290 mg, 524 μmol , 89%), mp 236 °C. ^1H NMR (500 MHz, CDCl_3): δ 2.74 (s, 6 H), 1.98 (s, 12 H) ppm. $^{13}\text{C}\{^1\text{H}\}$ NMR (126 MHz, CDCl_3): δ 139.9 (2 C), 137.2 (2 C), 134.9 (4 C), 126.2 (4 C), 25.9 (2 CH_3), 21.5 (4 CH_3) ppm. IR (ATR): 2996 (w), 2945 (m), 2917 (m), 1440 (m), 1364 (s), 1047 (m), 1023 (m), 958 (vs), 944 (s), 663 (s), 626 (m) cm^{-1} . MS (EI, 70 eV), m/z (%): 550 (10) [M^+], 390 (10), 375 (10), 313 (31), 202 (60), 156 (100), 101 (72). Anal. calcd. for $\text{C}_{18}\text{H}_{18}\text{Br}_4$ (553.96): C 39.03, H 3.28; found C 39.06, H 3.28.

3,3',5,5'-Tetrakis(isopropylthio)bimesitylene (12). Sodium isopropylthiolate (1.47 g, 15.0 mmol) was added to a solution of compound **13** (519 mg, 937 μmol) in DMA (15 mL) under inert atmosphere. The mixture was stirred at 100 °C for 20 h. After cooling to ambient temperature, water (80 mL) was added and the resulting suspension was extracted with *n*-hexane (3 x 50 mL). The combined organic extracts were washed with brine (50 mL), dried over MgSO_4 , filtered and the solvent was removed in vacuum. The residue was purified by column chromatography (SiO_2 , *n*-hexane/MTBE 10 : 1, R_f 0.81) to yield the title compound **12** as a colorless solid (500 mg, 935 μmol , 99%), mp 126 °C. ^1H NMR (500 MHz, CDCl_3): δ 3.09 (hept, J 6.6 Hz, 4 H), 2.93 (s, 6 H), 2.09 (s, 12 H), 1.19 (d, J 6.6 Hz, 24 H) ppm. $^{13}\text{C}\{^1\text{H}\}$ NMR (126 MHz, CDCl_3): δ 147.1 (2 C), 140.9 (4 C), 140.3 (2 C), 132.9 (4 C), 39.1 (4 CH), 23.0 (8 CH_3), 22.5 (2 CH_3), 19.8 (4 CH_3) ppm. IR (ATR): 2960 (s), 2922 (m), 2862 (m), 1441 (vs), 1375 (s), 1362 (m), 1243 (s), 1151 (m), 1052 (s), 1007 (m), 972 (m), 935 (m) cm^{-1} . MS (ESI, pos.), m/z : 557 [$\text{M} + \text{Na}^+$], 535 [$\text{M} + \text{H}^+$]. Anal. calcd. for $\text{C}_{30}\text{H}_{46}\text{S}_4$ (534.95): C 67.36, H 8.67, S 23.97; found C 67.37, H 8.66, S 23.98.

Bimesitylene-3,3',5,5'-tetrasulfonylchloride (14). *N*-Chlorosuccinimide (1.98 g, 14.8 mmol) was added portionwise to a suspension of compound **13** (496 mg, 927 μmol) in MeCN (11 mL) and hydrochloric acid (2.2 mL of a 2 mol/l solution in water, 4.4 mmol) at 0 °C. The reaction mixture stirred for 20 min. Water (20 mL) was then added and the aqueous layer was extracted with CH_2Cl_2 (3 x 10 mL). The combined organic layers were dried over MgSO_4 , filtered and the solvent was removed in vacuum. After purification by column chromatography (SiO_2 , CH_2Cl_2 , R_f 0.87) the product **14** could be obtained as a colorless solid (532 mg, 841 μmol , 91%), decomp. 170 °C. ^1H NMR (500 MHz, CDCl_3): δ 3.23 (s, 6 H), 2.39 (s, 12 H) ppm. $^{13}\text{C}\{^1\text{H}\}$ NMR (126 MHz, CDCl_3): δ 145.5 (4 C), 143.9 (4 C), 143.3 (2 C), 141.2 (2 C), 22.0 (2 CH_3), 20.6 (4 CH_3) ppm. IR (ATR): 2937 (w), 1391 (m), 1370 (s), 1345 (m), 1171 (vs), 701 (s), 624 (m), 543 (s) cm^{-1} . MS (EI, 70 eV), m/z (%): 630 (20) [M^+], 532 (70), 433 (45), 269 (73), 84 (76), 64 (100). Anal. calcd. for $\text{C}_{18}\text{H}_{18}\text{Cl}_4\text{O}_8\text{S}_4$ (632.37): C 34.19, H 2.87, S 20.28; found C 34.24, H 2.87, S 20.31.

Bimesitylene-3,3',5,5'-tetrasulfonic acid tetrahydrate, $\text{H}_4\text{BMTS} \cdot 4 \text{H}_2\text{O}$ (6). A suspension of compound **14** (136 mg, 215 μmol) in water (15 mL) was refluxed for 16 h. THF (8 mL) was then added and the solution stirred at reflux temperature for 6 h. After cooling to ambient temperature the reaction mixture was filtered and the filtrate concentrated in vacuum to yield the title compound **6** as a colorless, strongly hygroscopic solid (130 mg, 206 μmol , 96%). ^1H NMR (500

MHz, D₂O): δ 2.76 (s, 6 H), 2.00 (s, 12 H) ppm. ¹³C{¹H} NMR (126 MHz, D₂O): δ 142.0 (4 C), 141.3 (2 C), 137.6 (4 C), 136.6 (2 C), 21.3 (2 CH₃), 18.9 (4 CH₃) ppm. IR (ATR): 3367 (s br), 2937 (w), 1686 (m br), 1147 (s), 1109 (m), 1032 (s), 1017 (s), 671 (m), 526 (m) cm⁻¹. MS (ESI, neg.), *m/z*: 557 [M – H⁺], 278 [M – 2 H⁺]. HRMS (ESI, neg.): calcd. for C₁₈H₂₁O₂₂S₄⁻: 556.9916; found 556.9898 [M – H⁺].

Bis[pentaquacopper(II)] biphenyl-3,3',5,5'-tetrasulfonate dihydrate, [Cu(OH)₂]₅(BPTS) · 2 H₂O (15). Cu₂(OH)₂(CO₃) (33 mg, 0.15 mmol) was added to a solution of tetrasulfonic acid **4** (80 mg, 147 μ mol) in water (10 mL). The mixture was stirred at 50 °C for 20 h. The resulting suspension was filtered and the solvent evaporated to yield the title compound **15** as a blue solid (95 mg, 117 μ mol, 79%), which contained single crystals. ¹H NMR (500 MHz, D₂O): δ 8.35 (br, 6 H) ppm. ¹³C{¹H} NMR (126 MHz, D₂O): δ 144.5 (4 C), 140.7 (2 C), 127.3 (4 CH), 122.7 (2 CH) ppm. C₁₂H₆Cu₂O₁₂S₄ · 12 H₂O (813.68). The monoclinic crystal structure [*P*2₁/*c*, *Z* 2, *a* 5.8423(3), *b* 11.2646(6), *c* 20.995(1) Å, β 97.501°] was solved by direct methods and expanded using Fourier techniques. One of the water molecules in the structure was treated by a split model. The structure refinement using 6643 unique reflections led to residuals of R1 0.0275 and wR2 0.0720. CCDC-1013955 contains the supplementary crystallographic data for this compound. These data can be obtained free of charge via www.ccdc.cam.ac.uk, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12, Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

Bis[hexaquacopper(II)] bimesitylene-3,3',5,5'-tetrasulfonate, [Cu(OH)₂]₆(BMST) (16). Cu₂(OH)₂(CO₃) (93 mg, 0.42 mmol) was added to a solution of tetrasulfonic acid **6** (265 mg, 421 μ mol) in water (10 mL). The mixture was stirred at 50 °C for 20 h. The resulting suspension was filtered and the solvent evaporated to yield the title compound **16** as a blue solid (299 mg, 333 μ mol, 79%), which contained single crystals. C₁₈H₁₈Cu₂O₁₂S₄ · 12 H₂O (897.84). The tetragonal crystal structure [*I*-42*m*, *Z* 2, *a* 11.2232(7), *c* 13.119(1) Å] was solved by direct methods and expanded using Fourier techniques. Three carbon atoms of the phenyl rings of the anion show a slight disorder which could be easily resolved using a split model. The structure refinement based on 2128 unique reflections led to residuals of R1 0.0341 and wR2 0.0835. CCDC-1013956 contains the supplementary crystallographic data for this compound. These data can be obtained free of charge via www.ccdc.cam.ac.uk, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12, Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

Supplementary Material Available: ¹H and ¹³C NMR spectra of all reported compounds.

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