

Tailoring nostoclide structure to target the chloroplastic electron transport chain

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Dedicated to Prof. Pawel Kafarski to honor the achievements within his career

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Abstract

Aiming to improve their effectiveness, three modifications were introduced into the structure of the natural phytotoxins nostocclides, leading to the synthesis of novel 3-benzyl-4-isopropyl-5-(arylmethylene)furan-2(5*H*)-ones, 3-benzyl-5-(furan-2-ylmethylene)furan-2(5*H*)-ones, and 3,4-dihalo-5-arylidenefuran-2(5*H*)-ones. All compounds were characterized by IR, ¹H and ¹³C NMR, NOEDIF, COSY, HETCOR and MS spectrometry. Increasing the length of the molecule was found to reduce the ability to interfere with ferricyanide reduction by isolated chloroplasts. The addition of an isopropyl moiety into the heterocyclic ring, as in naturally-occurring nostocclides, did not influence the inhibitory potential. Also the replacement of the electron-drawing phenyl substituent with two halogen residues did not improve the resulting activity. However, both latter modifications influenced the interaction with the photosynthetic machinery. These analogues could therefore represent novel leads to be explored toward the development of new herbicides targeting the chloroplastic electron transport chain.

Keywords: Energy transfer inhibitors, herbicides, Hill reaction, nostoclide analogues, electron transport chain inhibitors, γ -alkylidenebutenolides

Introduction

Increasing public concern for the environmental pollution deriving from agricultural practice strictly requires that modern agrochemicals would be endowed with low persistence. Intensive efforts have been undertaken to discover new compounds with favorable environmental and

safety features to selectively control weeds.^{1,2} The identification of novel herbicides is highly desirable also to overcome weed resistance, rapidly raised as a consequence of the severe selective pressure imposed by a continuous application of active ingredients with the same biochemical mode of action.³ However, the unfavorable economics of herbicide development, mainly due to increasing regulatory requirements, is greatly limiting industrial discovery programs.⁴ Moreover, many active principles that are currently in use came from a small set of molecular scaffolds, whose functional lifetime has been extended by synthetic tailoring (*e.g.* triazines, sulfonyleureas and imidazolinones).⁵ Because of the exponential emergence and diffusion of herbicide resistance, the discovery of new scaffolds is mandatory. The exploitation of natural products might combine the need of new chemical structures with a suitable susceptibility to microbial degradation.⁶

Among the naturally-occurring phytotoxins discovered to date are the nostocliides (Figure 1), γ -alkylidenebutenolides produced by *Nostoc sp.*⁷ The nostocliides resemble cyanobacterin, a lactone isolated from the blue-green alga *Scytonema hofmanni*, which inhibits the growth of several eukaryotic algae and mono- and dicotyledonous angiosperms by interfering with the photosynthetic electron transport.⁸

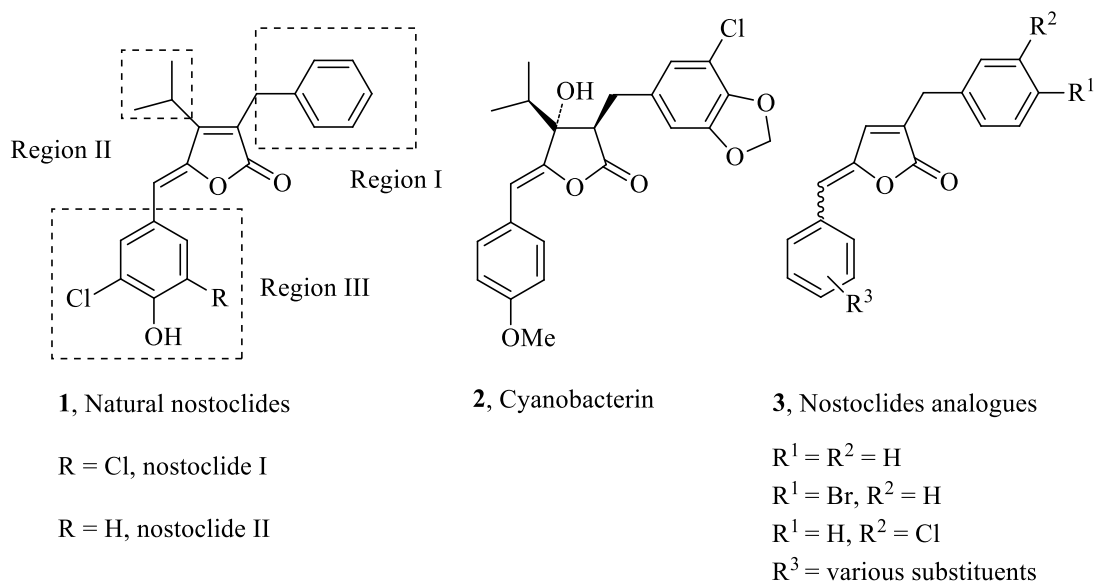


Figure 1. Naturally-occurring phytotoxins and chemically synthesized analogues. **1:** nostocliides; **2:** cyanobacterin; **3:** 3-benzyl-5-(arylmethylene)furan-2(5*H*)-ones designed using nostocliides as a lead structure.^{9,10}

When a group of twelve 3-(4-bromobenzyl)-5-(arylmethylene)furan-2(5*H*)-ones were synthesized using the nostocliides as a lead structure, several of them were indeed found to inhibit the light-driven reduction of ferricyanide by isolated spinach chloroplasts, the so-called Hill reaction.⁹ A second series of 34 non-brominated analogues showed better inhibitory properties,

but none of them was capable of completely suppressing the Hill reaction, mostly due to a poor water solubility.¹⁰ The inhibition brought about by the most active brominated analogues was found to be progressively reversed when increasing NH_4Cl concentrations dissipated the proton gradient across the thylakoid membrane.⁹ These results suggest that the synthetic nostoclide analogues may act by inhibiting phosphorylation in chloroplasts, working as energy transfer inhibitors through an interaction with the coupling factor, $\text{CF}_0\text{-CF}_1$. This mode of action differs from that of commercial herbicides targeting photosynthesis, most of which interact with the electron transport chain,¹¹ but is similar to that reported for other naturally-occurring phytotoxins.¹² A thorough quantitative structure-activity relationship analysis suggested that the most relevant molecular property modulating the inhibitory potential of this class of compounds is their molecular polarity, as expressed by the dipole moment.¹³

A certain degree of lipophilicity is usually required for a photosynthetic inhibitor, given that in order to exert its effect, it needs to diffuse through both the plasmalemma and the chloroplast envelopes.¹⁴ In several cases the presence of a hydrocarbon side chain has been reported to greatly impact the activity of a given scaffold, as it provides the required lipophilicity.¹⁵ But the opposite is also true, because sufficient hydrophilicity is needed to allow herbicides to be translocated within the plants and reach their target inside the chloroplast.

To obtain more effective derivatives, some modifications were designed into the scaffold of nostoclide. As shown in Figure 1, their γ -alkylidenebutenolide core has three substituent groups (Regions I, II, and III). In this work we report the synthesis of compounds where these regions were modified in order to evaluate the effect of each group on the resulting ability to interfere with the photosynthetic apparatus.

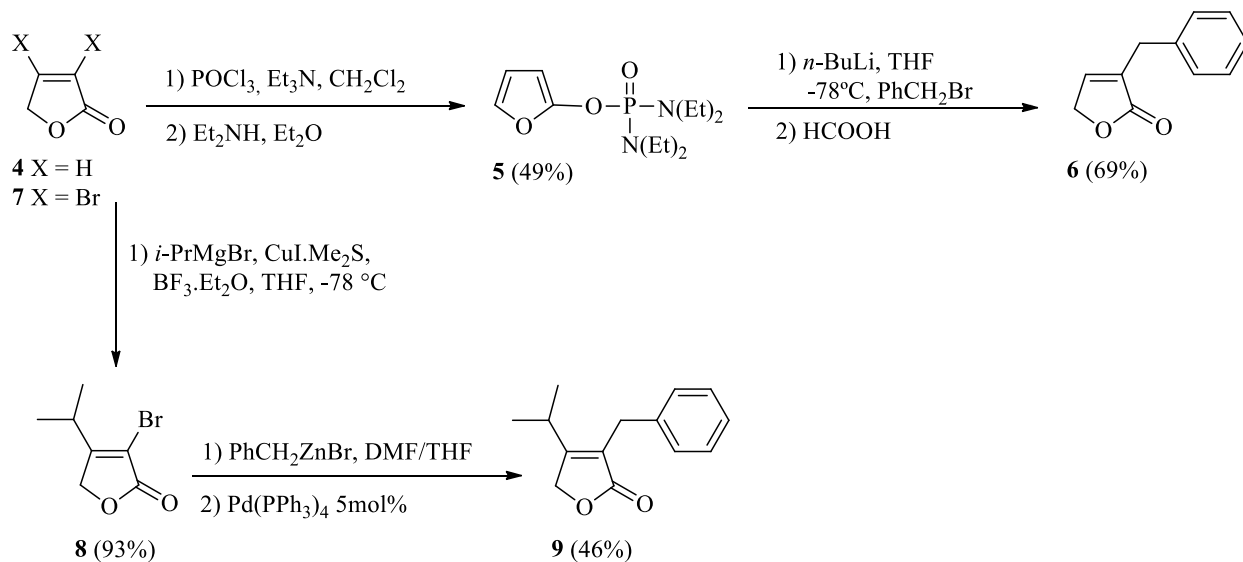
Results and Discussion

Synthesis

Compound **4**, used as starting material for the synthesis of **6**, is commercially available, but due to its high price (US\$200/gram) it was prepared from the cheap furfuraldehyde by oxidation with hydrogen peroxide followed by basic hydrolysis.¹⁶ Compound **4** was then converted into the phosphorous intermediate **5** in 49% yield using a methodology previously described.⁹ Further treatment of **5** with BuLi followed by capture of the furyllithium formed with benzylbromide and acid hydrolysis of the intermediate, resulted in the formation of **6** in 69% yield (Scheme 1). This compound is not new and as a consequence the full spectroscopic data are not reported, as they were in full agreement with previously reported data.^{10, 17}

In order to evaluate the influence of the isopropyl group on the biological activity, the intermediate **9** was prepared as shown in Scheme 1. Compound **9** has been previously prepared by Bellina and Rossi.¹⁸ They introduced the isopropyl group by reaction of **7** with toxic prop-2-enyltributyltin, followed by the hydrogenation of the propenyl unit using expensive Wilkinson catalyst. This two-step procedure resulted in compound **8** in 74% yield. Further palladium

coupling of the benzyl group resulted in **9**. As the overall yield obtained was only 26.6%, we developed an alternative procedure for the synthesis of **9**. In our synthesis, compound **7** was reacted with isopropyl magnesium bromide in the presence of CuI or CuBr.Me₂S and using the Lewis acids BF₃.Et₂O or TMSCl as a catalyst and THF or diethyl ether as a solvent. The reaction conditions were optimized and compound **8** was produced in one step from **7** in 93% yield upon the use of CuBr.Me₂S (1 eq) and BF₃.Et₂O (2 eq). The best solvent for this reaction was THF, whereas diethyl ether did not afford the required product in good yield.



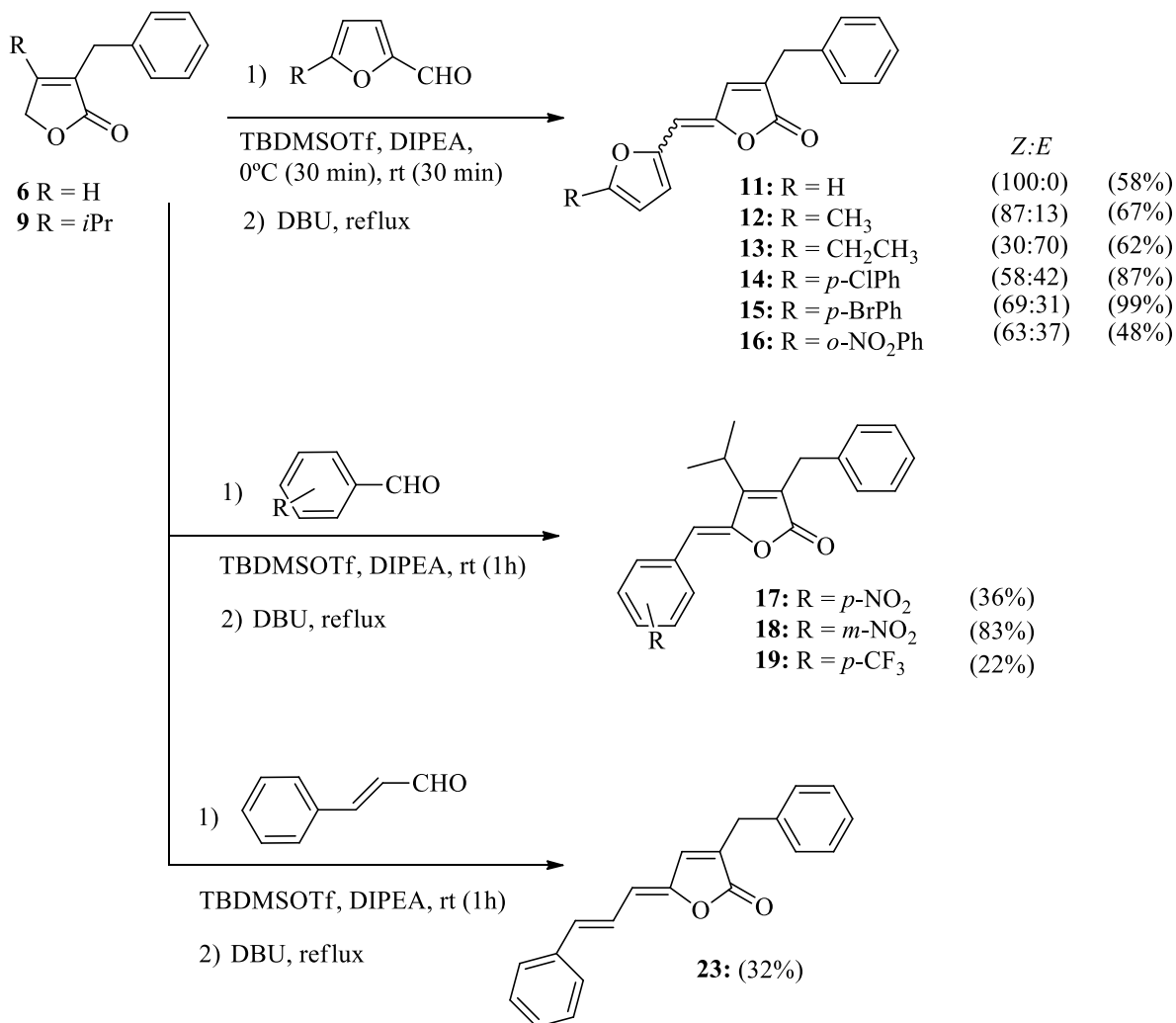
Scheme 1. Synthesis of precursor compounds **6** and **9**.

Further introduction of the benzyl group at the alpha position was possible by means of the Negishi coupling¹⁹ and the best experimental conditions led to the required product in 46% yield. The spectroscopic data (not shown) obtained for **9** were in full agreement with those previously reported.^{17, 18}

With compound **6** and **9** in hands, new nostoclide analogues were prepared (Scheme 2). The first series prepared consisted in compounds **11-16**, all containing a furanmethylene unit attached to the lactone ring. The rationale for the introduction of the furan ring was based on the results previously reported that the biological activity was associated with the polarity of the compounds.¹³ We anticipated that as furyl unit is more polar than the phenyl ring, the introduction of this heteroaromatic moiety could result in more active substances.

The appendage of the C5 substituents at compounds **6** and **9** can be achieved in two steps by aldol condensation followed by elimination of water²⁰ or by a one-pot arylmethylenation procedure developed by Boukouvalas^{17,21} and several procedures were reported for the aldol reaction involving butenolide **4**. Initially, we tried the new condensation protocol described by Xu et al.,²² but contrary to what they reported, only one product was isolated from the hydrolysis of the butenolide. Then reaction of **6** with furfural in the presence of *tert*-butyldimethyl-

silyltrifluoro-methanesulfonate, and diisopropyl-ethylamine followed by treatment of the silyl ether generated *in situ* with DBU¹⁷, led to the formation of compound **11** in 58% yield. In this case only the *Z* isomer was isolated as it was observed for the preparation of its benzylidene analogues **24-32** (Scheme 2). The confirmation of the structure of **11** was based on its spectroscopic data, including Nuclear Overhauser Effect difference spectroscopy (NOEdiff) that revealed a correlation between H-6 and H-4, with 6% increase of the signal of H-4, in agreement with the *Z* stereochemistry.



Scheme 2. Synthesis of compounds **11-19** and **23**.

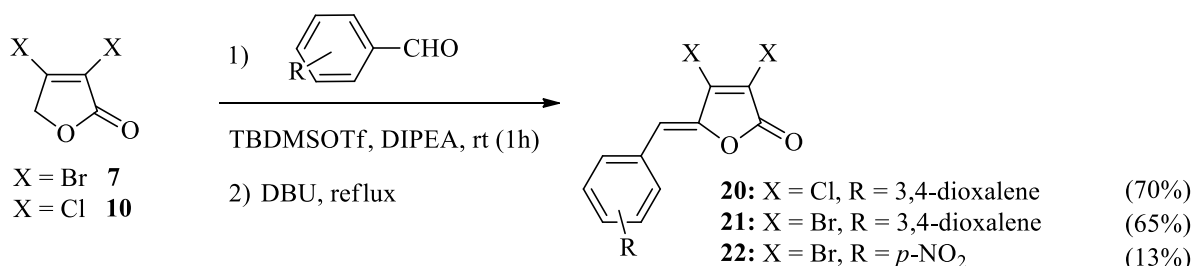
Compounds **12** to **16** were prepared using the same condensation reaction conditions but in such cases mixtures of *Z/E* isomers were formed. The ratio of *Z/E* varied from 87:13 in the case of **12** to 30:70 for compound **13**. All isomers were separated by HPLC analysis and fully characterized, except the *E* isomer of **12**, as it degraded during the purification steps. For all *Z*

isomers a NOE effect was observed between H-6 and H-4 as in the case of compound **11**. In the case of the *E* isomers, no NOE effect was observed.

In a previous work we have shown that nostoclides analogues with electron withdrawing groups in the benzylidene rings were among the most active.^{9,10} So, to evaluate the influence of the isopropyl group, present in the natural nostoclides, compound **9** was converted into benzylidene lactones **17** (36%), **18** (83%), and **19** (22%). A few attempts to improve the yields of **17** and **19** were carried out, but at this point of the work, we concentrated our effort in the production of enough material for biological investigation. All compounds were produced as *Z* isomer, being that a nuclear NOE effect was observed between H-6 and hydrogen of group isopropyl, and even for the low yield reactions.

Compound **23** was prepared from 3-phenylpropenal in 32% yield under the same conditions used before. No attempt was made to optimize the yield of this product as our major goal was to evaluate the influence of the extra C=C carbon unit on the biological activity as this has not been investigated in previous works.^{9,10}

To evaluate if the benzyl unit at carbon 3 of the furanone ring is essential for the biological activity, we prepared three halogenated analogues **20-22**. They were synthesized from 3,4-dibromofuran-2(5*H*)-one **7** and 3,4-dichlorofuran-2(5*H*)-one **10** in 13% to 65% yields, as depicted in Scheme 3, using the same alkylidenation protocol already described. Although compound **22** was prepared from the activated electron poor aldehyde, it was isolated in very small yield. This result is not easily explained considering that the more electron-rich aldehyde 3,4-dioxalenebenzaldehyde, resulted in high yield condensation products **20** and **21**.²⁰



Scheme 3. Synthesis of compounds **20-22**.

The stereochemistry of compounds **20-22** is *Z* since the *E* isomer would be formed via an intermediate less stable due to the repulsion between the bromine at beta position and the aryl group.²¹

Biological activity

The ability of these three small groups of nostoclides analogues to interfere with the light-driven ferricyanide reduction by osmotically swollen spinach chloroplasts was then investigated. The results were compared with those obtained with the corresponding 3-benzyl-5-(arylmethylene)

furan-2(5*H*)-one counterparts (Figure 2), whose biological activity had been previously characterized.^{9,10}

Data concerning compounds **14**, **15** and **23** (Table 1) clearly suggest that the increase of the length of the molecule does not increase the inhibitory potential. Indeed, the inhibition brought about by poorly active nostoclide was not significantly improved, whereas that of an active derivative lowered as a consequence. On this basis, the higher effectiveness found for several derivatives bearing bulky substituents^{9,10} has most likely to be interpreted as a consequence of the resulting effect on the electronic distribution, and not of an insufficient spanning of the scaffold across the thylakoid membrane.

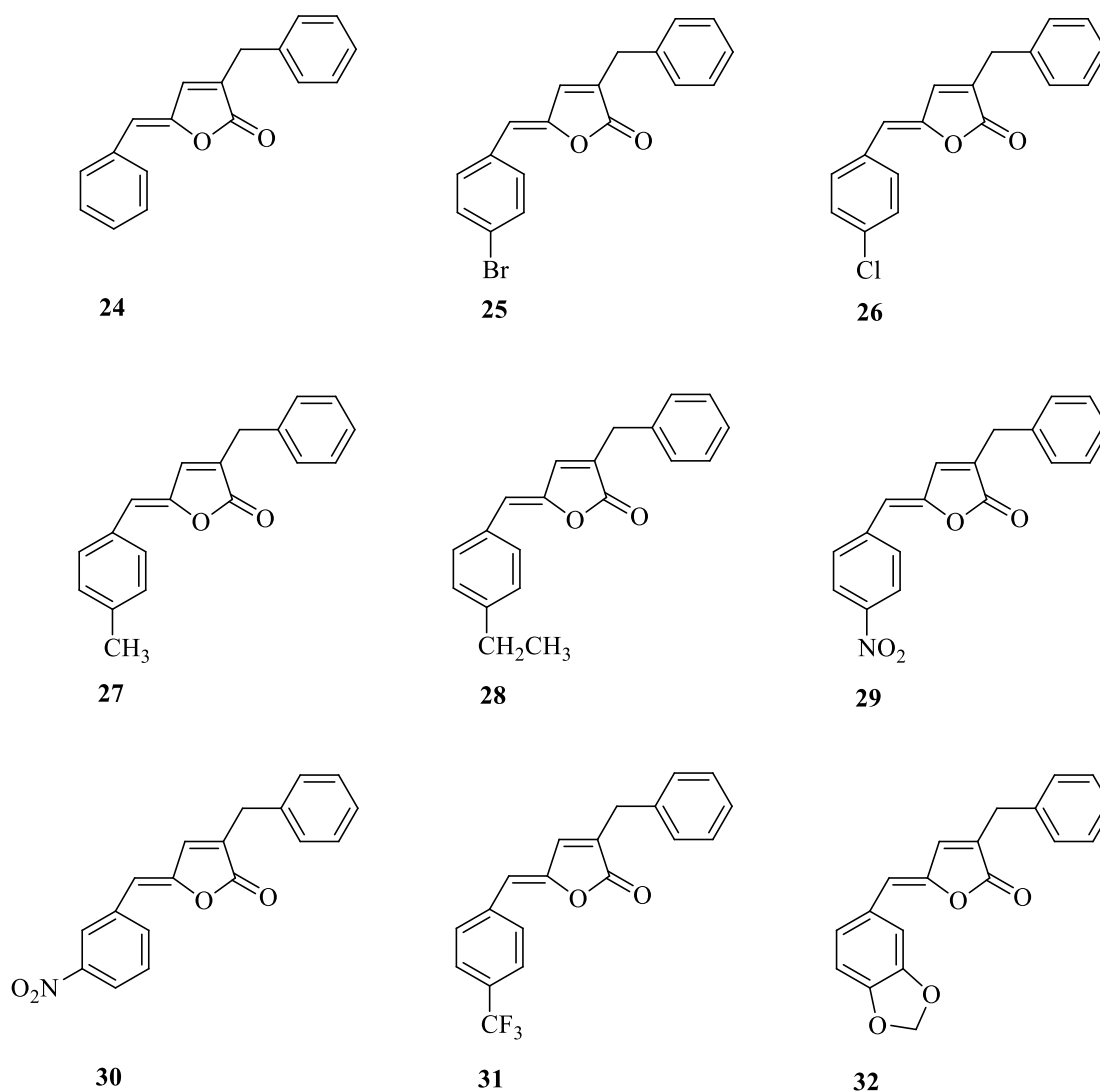


Figure 2. Nostoclide analogues previously synthesized and tested for the ability to interfere with the photosynthetic electron transport chain.^{9,10} These compounds were included for comparison.

Table 1. Concentrations able to 50%-inhibit the photosynthetic electron transport chain, measured as the ability to reduce ferricyanide by functionally intact spinach chloroplasts

Compound	IC ₅₀ (μM)	Nostoclides analogue	IC ₅₀ (μM)
14	380 ± 27	26	>1000
15	>1000	25	473 ± 42
23	118 ± 23	24	19.7 ± 3.9
17	2.9 ± 0.7	29	1.7 ± 0.7
18	190 ± 23	30	23.1 ± 13.3
19	15.8 ± 5.1	31	8.3 ± 2.3
20	>1000	32	>1000
21	>1000	32	>1000
22	6.2 ± 0.6	29	1.7 ± 0.7

Data are compared with those obtained with the corresponding nostoclides analogues previously synthesized and evaluated as Hill reaction inhibitors.^{9,10} Data presented for compounds **14** and **15** refer to the *E* isomers; the experiment was performed also with the *Z* isomers, obtaining very similar patterns (data not shown).

With respect to compounds **17-19**, the presence of an isopropyl moiety in the heterocyclic ring, as in natural nostoclides and cyanobacterin, did not exceedingly vary the resulting ability to inhibit ferricyanide reduction (Table 1). This allows us to rule out the possibility that the remarkably higher inhibition brought about by the latter natural toxin⁸ may be due to this feature. The result seems also consistent with previous data showing that any change of the substituents of the phenyl ring in the cyanobacterin molecule produced less phytotoxic analogues, even if those of the furan ring remained unchanged.²³

Also in the case of compounds **20-22** the modifications introduced in the scaffold failed to increase the inhibitory activity. On the contrary, a slight decrease was evident for compound **22** (Table 1). However, this was paralleled by a concomitant increase of its water solubility, one that allows achieving actually high concentrations in the reaction mixture (not shown). Over a certain threshold, at about 2 μM, the low solubility of the 3-benzyl-5-(arylmethylene)furan-2(5*H*)-one counterpart indeed caused the resulting inhibitory effect to rapidly reach a plateau and not increase further. The lack of a dose-effect relationship for the active compounds at high rates is the reason why the nostoclides analogues previously synthesized were unable to completely suppress the Hill reaction.^{9,10} On the contrary, over the whole range tested compound **22** was found to progressively inhibit ferricyanide reduction, which at concentrations exceeding 20 μM was almost undetectable.

This preliminary characterization was performed by measuring ferricyanide reduction under basal conditions. To gain more information about their mode of action, the inhibitory properties of the two most active compounds **17** and **22** were assessed also under phosphorylating or uncoupled conditions. In the presence of substrates of ATP synthase, both compounds retained their ability to inhibit the Hill reaction, and the residual percent activity did not differ significantly from that evaluated under basal conditions. Unexpectedly, the same result was found with the uncoupled electron flow (data not presented).

This behavior differs from that previously found for brominated 3-benzyl-5-(arylmethylene)furan-2(5*H*)-one lactones, whose effect was progressively reverted when increasing levels of an uncoupler were included into the reaction mixture.⁹ On this basis, those nostoclide analogues were supposed to be able to interfere with phosphorylation in chloroplasts, acting as energy transfer inhibitors through an interaction with the coupling factor, CF₀-CF₁. To verify this aspect, the inhibition brought about by compounds **17** and **22** was also assessed at varying ammonia concentration. As a term of comparison, two 3-benzyl-5-(arylmethylene)furan-2(5*H*)-ones (either brominated or not) also bearing a nitroso moiety in the phenyl ring were included in the experimental plan. Results (Figure 3) confirmed a difference among these compounds.

Consistently with previous findings, the effect of the brominated 3-benzyl-5-(arylmethylene)furan-2(5*H*)-one was completely relieved when the concentration of the uncoupler was high enough to dissipate the transmembrane proton gradient. Interestingly, the inhibition brought about by the non-brominated analogue was reversed only in part, whereas that of both the new compounds was insensitive at all to the presence of ammonia. These data therefore suggest that, at varying the chemical structure of nostoclide scaffold, the interaction with the photosynthetic machinery also changes. The brominated analogue most likely interacts only with the coupling factor. The non-brominated derivative seems to interact both with the coupling factor and the electron transport chain, since it is still effective under uncoupling conditions, although to a lower extent. Compounds **17** and **22** may on the contrary interfere only with the electron transport chain, being their inhibitory activity substantially unchanged under all the experimental conditions employed. Because this mode of action is the same of commercial herbicides targeting photosynthesis,¹¹ their effectiveness *in vivo* could be higher than that of the 3-benzyl-5-(arylmethylene)furan-2(5*H*)-one lactones previously synthesized. Experiments are currently in progress to verify this possibility.

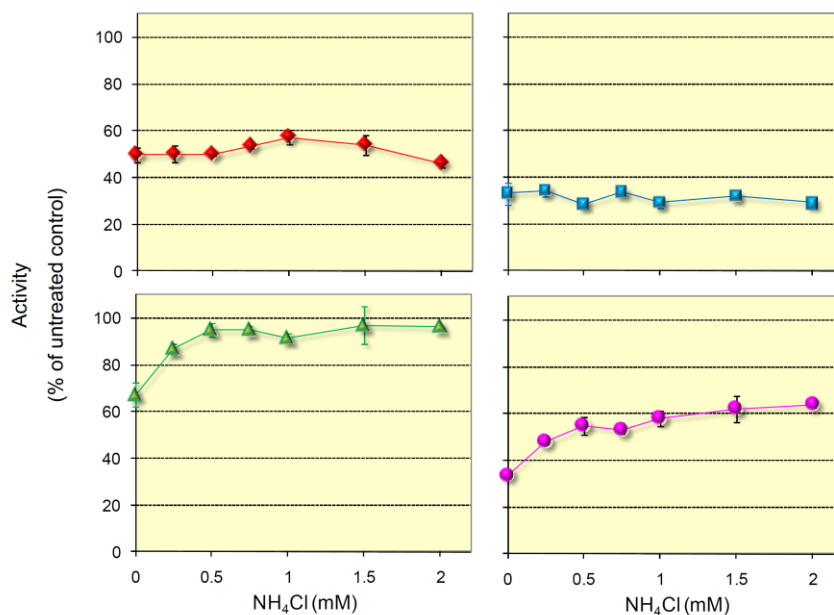


Figure 3. Effect of a fixed concentration of compounds **17** (5 μM \blacklozenge) and **22** (10 μM \blacksquare) on the photosynthetic electron transport chain at increasing ammonia concentration. Data are shown together with those obtained with the corresponding brominated (10 μM compound **30** \blacktriangle) or non-brominated (2.5 μM compound **29** \bullet) 3-benzyl-5-(arylmethylene)furan-2(5*H*)-one lactones also bearing a nitroso moiety in the phenyl ring. Results are expressed as percentage of those obtained with the same NH_4Cl concentration in the absence of any inhibitor.

Experimental Section

General. Solvents and amines were purified as described by Perrin and Armarego.²⁷ All other reagents were purchased from Sigma Aldrich (Milwaukee, WI, USA) and utilized without further purification. Lactone **4** was synthesized with 44% yield, from furfural, as previously described.¹⁰ The compounds **7** and **10** were prepared in 75% and 60% yield, respectively from mucohalic acid according to the literature²⁴ and lactones **6** and **9** were synthesized as previously described.^{10,17} Commercially available *n*-butyllithium hexane solutions (1.4 mol L⁻¹) were titrated prior to use.²⁵ Melting points are uncorrected and were obtained from an MQAPF-301 melting point apparatus (Microquimica, Brazil). Analytical thin layer chromatography analysis was conducted on aluminum-packed precoated silica gel plates. Column chromatography was performed over silica gel (70-230 or 230-400 mesh). Infrared spectra were recorded on a Perkin-Elmer Paragon 1000 FTIR spectrophotometer or Varian 660-IR FTIR with accessory PIKE Gladi ATR, using potassium bromide (1% w/w) disks, or thin liquid film on NaCl discs. Mass spectra were recorded on a Shimadzu GCMS-QP5050A instrument by direct insertion, EI mode (70 eV). The ¹H and ¹³C NMR spectra were recorded on a Varian Mercury 300 spectrometer at 300 and

75 MHz, respectively, using CDCl_3 as solvent and TMS as internal standard, unless otherwise stated.

The HPLC purifications were performed on a Shimadzu LC-10A equipment, using an UV detector at 292 nm and a preparative column Shim-Pack 20 x 250 mm (15 μm). The mobile phase for all the purifications was dichloromethane and hexane and flow of 6 mL/min and the injection volume was 200 μL .

(Z)-3-Benzyl-5-(furan-2-ylmethylene)furan-2(5H)-one (11). To a two-neck round-bottom flask under nitrogen atmosphere were added 3-benzylfuran-2(5H)-one **6** (100 mg, 0.57 mmol), dichloromethane (DCM, 3 mL), TBDMSOTf (170 μL , 0.74 mmol), DIPEA (310 μL ; 1.2 mmol) and 2-furaldehyde (0.12 g; 1.2 mmol). The reaction mixture was stirred for 30 min at 0 °C and a further 30 min at room temperature. After DBU (120 μL , 1.2 mmol) was added, the reaction mixture was refluxed for an additional 1h before the addition of dichloromethane (70 mL). The resulting organic layer was washed with 3 mol L^{-1} HCl aqueous solution (2×25 mL) and brine (25 mL). After separation, the organic layer was dried over MgSO_4 , filtered, and concentrated under reduced pressure. The crude residue was purified by flash column chromatography on silica-gel using DCM /hexane (2:1 v/v) as elutant to yield the required product as an yellow solid in 58% yield. Mp 76.1-77.0 °C, IR (KBr) 3086, 3062, 3028, 2922, 1763, 1650, 1603, 1560, 1474, 1335, 1272, 1078 cm^{-1} . ^1H NMR (300 MHz, CDCl_3), δ 3.72 (s, 2H, H-7), 5.91 (m, 1H, H-6), 6.52 (ddd, 1H, H-4''); $J_{4''/3''}$ 3.4 Hz, $J_{4''/5''}$ 1.8 Hz, $J_{4''/6}$ 0.6 Hz), 6.92 (td, 1H, H-4; $J_{4/7}$ 1.5 Hz, $J_{4/6}$ 0.3 Hz), 6.99 (dl, 1H, H-3''; $J_{3''/4''}$ 3.4 Hz), 7.24-7.38 (m, 5H, Ph), 7.47 (dd, 1H, H-5''; $J_{5''/4''}$ 1.8 Hz, $J_{5''/3''}$ 0.9 Hz). ^{13}C NMR (75 MHz, CDCl_3), δ 32.0 (C-7), 101.2 (C-6), 113.2 (C-4''), 114.9 (C-3''), 127.2 (C-4'), 129.1 (C-3'/5'), 129.1 (C-2'/6'), 132.9 (C-3), 137.5 (C-1'), 138.5 (C-4), 144.0 (C-5''), 145.7 (C-5), 149.3 (C-2''), 170.2 (C-2). HRMS (ESI) Calcd for $\text{C}_{16}\text{H}_{12}\text{NaO}_3$ (M+Na), 275.0679; Found, 275.0678.

Compounds 12-16 were prepared by employing a procedure similar to that described for compound **11**. Compounds **17-23** were prepared as previously described.^{9,10} The synthesized compounds were fully characterized by IR, NMR (^1H and ^{13}C), COSY, HETCOR, NOEDIF and MS spectrometry. Structures for the remaining compounds are supported by the following spectroscopic data.

(Z)-3-Benzyl-5-(5-methylfuran-2-ylmethylene)furan-2(5H)-one (12a). Yield 67% (Z:E 87:13), diastereoisomers obtained after purification by column chromatography, eluent DCM/hexane (2:1 v/v). Yellow solid, mp 87.4-87.7 °C, isolated by HPLC, eluent DCM /hexane (1:1 v/v). IR (neat) 3062, 3028, 2921, 1761, 1650, 1584, 1517, 1495, 1429, 1381, 1276, 1129, 1075, 1024 cm^{-1} . ^1H NMR (300 MHz, CDCl_3), δ 2.33 (s, 3H, CH_3), 3.71 (s, 2H, H-7), 5.84 (s, 1H; H-6), 6.12-6.14 (m, 1H, H-4''), 6.89-6.90 (m, 2H, H-4, H-3''), 7.24-7.35 (m, 5H, Ph). ^{13}C NMR (75 MHz, CDCl_3), δ 14.1 (CH_3), 32.0 (C-7), 101.5 (C-6), 109.9 (C-4''), 116.6 (C-3''), 127.1 (C-4'),

129.0 (C-3'/5'), 129.1 (C-2'/6'), 131.9 (C-3), 137.6 (C-1'), 138.5 (C-4), 144.8 (C-5), 147.9 (C-2''), 154.7 (C-5''), 170.0 (C-2). HRMS (ESI) Calcd for C₁₇H₁₄NaO₃ (M+Na), 289.0835; Found, 289.0827.

(E)-3-Benzyl-5-(5-methylfuran-2-ylmethylene)furan-2(5H)-one (12b). The compound degraded in the presence of silica during isolation by HPLC, using eluent DCM/hexane (1:1 v/v) is not possible to obtain all spectra. ¹H NMR (300 MHz, C₆D₆), δ 1.69 (s, 3H, CH₃), 3.41 (s, 2H, H-7), 5.54 (br d, 1H, H-4''); *J*_{4''/3''} 3.0 Hz), 5.79 (br dl, 1H, H-3''); *J*_{3''/4''} 3.0 Hz), 5.90 (brd., 1H, H-6), 7.58-7.61 (m, Ph, H-4). HRMS (ESI) Calcd for C₁₇H₁₄NaO₃ (M+Na), 289.0835; Found, 289.0827.

(Z)-3-Benzyl-5-(5-ethylfuran-2-ylmethylene)furan-2(5H)-one (13a). Yield 62% (*Z:E* 30:70), diastereomers obtained after purification by column chromatography, eluent DCM /hexane (1:1 v/v). Orange amorphous solid, isolated by HPLC, eluent DCM /hexane (1:1 v/v). IR (neat) 3062, 3029, 2973, 2937, 1755, 1647, 1514, 1494, 1014 cm⁻¹. ¹H NMR (300 MHz, C₆D₆), δ 0.97 (t, 3H, CH₃; *J*_{CH₃/CH₂} 7.5 Hz), 2.32 (q, 2H, CH₂; *J*_{CH₂/CH₃} 7.5 Hz), 3.33 (s, 2H, H-7), 5.41 (s, 1H, H-6), 5.79-5.81 (m, 1H, H-4''), 6.02-6.03 (m, 1H, H-4), 6.89-7.12 (m, 6H, Ph, H-3''). ¹³C NMR (75 MHz, C₆D₆), δ 11.8 (CH₃), 21.6 (CH₂), 31.8 (C-7), 100.5 (C-6), 108.3 (C-4''), 115.89 (C-3'/5'), 127.57 (C-4'), 128.8 (C-2'/6'), 128.2-128.8 (C-3), 129.1 (C-3''), 131.9 (C-1'), 137.8 (C-4), 145.3 (C-5), 148.1 (C-2''), 159.4 (C-5''), 169.4 (C-2). HRMS (ESI) Calcd for C₁₈H₁₆NaO₃ (M+Na), 303.0992; Found, 303.0983.

(E)-3-Benzyl-5-(5-ethylfuran-2-ylmethylene)furan-2(5H)-one (13b). Yield 62% (*Z:E* 30:70), diastereoisomers obtained after purification by column chromatography, eluent DCM /hexane (1:1 v/v). Orange amorphous solid; isolated by HPLC, eluent DCM /hexane (1:1 v/v). IR (KBr) 3062, 3029, 2974, 2937, 1744, 1600, 1495, 1454, 1319, 1265, 1015 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 1.19 (t, 3H, CH₃; *J*_{CH₃/CH₂} 7.6 Hz), 2.65 (q, 2H, CH₂; *J*_{CH₂/CH₃} 7.6 Hz), 3.75 (s, 2H, H-7), 6.03 (dt, 1H, H-4''); *J*_{4''/3''} 3.3 Hz; *J*_{4''/CH₂} 0.9 Hz), 6.25 (br s, 1H; H-6), 6.38 (d, 1H, H-3''); *J*_{3''/4''} 3.3 Hz), 7.24-7.37 (m, 5H, Ph), 7.65-7.66 (m, 1H, H-4). ¹³C NMR (75 MHz, CDCl₃), δ 12.1 (CH₃), 21.8 (CH₂), 32.2 (C-7), 102.1 (C-6), 107.4 (C-4''), 116.0 (C-3''), 127.1 (C-4'), 129.0 (C-3'/5'), 129.1 (C-2'/6'), 134.0 (C-3), 137.0 (C-4), 137.6 (C-1'), 146.6 (C-5), 148.0 (C-2''), 160.6 (C-5''), 170.0 (C-2). HRMS (ESI) Calcd for C₁₈H₁₆NaO₃ (M+Na), 303.0989; Found, 303.0983.

(Z)-3-Benzyl-5-((5-(4-chlorophenyl)furan-2-yl)methylene)furan-2(5H)-one (14a). Yield 87% (*Z:E* 58:42), diastereoisomers obtained after purification by column chromatography, eluent DCM /hexane (2:3 v/v). Yellow solid, mp 156.5-158.1 °C, isolated by HPLC, eluent DCM /hexane (2:1 v/v) and recrystallization with a mixture of DCM /hexane. IR (KBr) 3098, 3032, 2922, 1749, 1641, 1474, 1452, 1373, 1283, 1036 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 3.73 (s, 2H, H-7), 5.95 (s, 1H, H-6), 6.78 (d, 1H, H-4''); *J*_{4''/3''} 3.6 Hz), 6.93 (m, 1H, H-4), 7.07 (d, 1H, H-3''); *J*_{3''/4''} 3.6 Hz), 7.25-7.38 (m, 7H, H-3''/5''; H-2'/6', H-3'/5', H-4'), 7.61-7.64 (m, 2H, H-2''/6''). ¹³C NMR (75 MHz, CDCl₃), δ 32.1 (C-7), 100.9 (C-6), 109.3 (C-3''), 117.4 (C-4''), 125.5 (C-2''/6''), 127.2 (C-4'), 128.7 (C-3''/5''), 129.1 (C-3'/5'), 129.2 (C-2'/6'), 129.3 (C-1''), 132.7 (C-3), 134.0 (C-1'), 137.4 (C-4''), 138.3 (C-4), 146.0 (C-5), 149.2 (C-2''), 154.1 (C-5''), 170.0 (C-2). HRMS (ESI) Calcd for C₂₂H₁₅ClNaO₃ (M+Na), 385.0602; Found, 385.0592.

(E)-3-Benzyl-5-((5-(4-chlorophenyl)furan-2-yl)methylene)furan-2(5H)-one (14b). Yellow solid, mp 202.3-204.0 °C, purified by HPLC, eluent DCM /hexane (2:1 v/v) and recrystallization with a mixture of DCM /hexane. IR (KBr) 3059, 2925, 1730, 1597, 1473, 1258, 1150, 1030 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 3.78 (s, 2H, H-7), 6.29 (s, 1H, H-6), 6.53 (d, 1H, H-3''; *J*_{3''/4''} 3.6 Hz), 6.66 (d, 1H, H-4''; *J*_{4''/3''} 3.6 Hz), 7.32-7.44 (m, 9H, H-2'''/6''', H-3'''/5''', H-2'/6', H-3'/5', H-4'). ¹³C NMR (75 MHz, CDCl₃), δ 32.5 (C-7); 101.2 (C-6), 108.5 (C-4''), 116.9 (C-3'''), 125.3 (C-2'''/6'''), 127.72 (C-4'), 129.1 (C-3'''/5'''), 129.3 (C-3'/5'), 129.4 (C-2'/6'), 132.7 (C-1'''), 135.6 (C-1'''), 137.3 (C-1'), 137.4 (C-3), 136.8 (C-4), 147.7 (C-5), 149.6 (C-2''), 154.7 (C-5''), 169.4 (C-2). HRMS (ESI) Calcd for C₂₂H₁₅ClNaO₃ (M+Na), 385.0602; Found, 385.0590.

(Z)-3-Benzyl-5-((5-(4-bromophenyl)furan-2-yl)methylene)furan-2(5H)-one (15a). Yield 99% (Z:E 69:31), diastereoisomers obtained after purification by column chromatography, eluent DCM/hexane (1:1 v/v). Yellow solid, mp 172.8-174.9 °C, isolated by HPLC, eluent DCM /hexane (1:1 v/v) and recrystallization with a mixture of DCM /hexane. IR (KBr) 3029, 2922, 2854, 1741, 1599, 1471, 1028 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 3.73 (s, 2H, H-7), 5.94 (s, 1H, H-6), 6.79 (d, 1H, H-4''; *J*_{4''/3''} 3.6 Hz), 6.93 (t, 1H; H-4; *J*_{4/7} 1.2 Hz), 7.07 (d, 1H; H-3''; *J*_{3''/4''} 3.6 Hz), 7.25-7.58 (m, 9H, H-2'''/6''', H-3'''/5''', H-2'/6', H-3'/5', H-4'). ¹³C NMR (75 MHz, CDCl₃), δ 32.1 (C-7), 100.9 (C-6), 109.4 (C-4''), 117.4 (C-3'''), 122.15 (C-4'''), 125.7 (C-2'''/6'''), 127.2 (C-4'), 129.0 (C-3'''/5'''), 129.1 (C-3'/5'), 132.2 (C-2'/6'), 132.7 (C-1'''), 137.3 (C-3), 137.4 (C-1'), 138.3 (C-4), 146.1 (C-5), 149.2 (C-2''), 154.2 (C-5''), 170.1 (C-2). HRMS (ESI) Calcd for C₂₂H₁₅BrNaO₃ (M+Na), 431.0082; Found, 431.0083.

(E)-3-Benzyl-5-((5-(4-bromophenyl)furan-2-yl)methylene)furan-2(5H)-one (15b). Yellow solid, mp 204.5-206.9 °C, purified by HPLC, eluent DCM/hexane (1:1 v/v) and recrystallization with a mixture of DCM/hexane. IR (KBr) 3060, 3028, 1743, 1645, 1600, 1471, 1238, 1073, 1036 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 3.78 (s, 2H, H-7), 6.28 (s, 1H, H-6), 6.52 (d, 1H, H-3''; *J*_{4''/3''} 3.6 Hz), 6.67 (d, 1H, H-4''; *J*_{4''/3''} 3.6 Hz), 7.24-7.49 (m, 9H, H-2'''/6''', H-3'''/5''', H-2'/6', H-3'/5', H-4'), 7.60-7.61 (m, 1H, H-4). ¹³C NMR (75 MHz, CDCl₃), δ 32.5 (C-7), 101.8 (C-6), 108.5 (C-4''), 116.9 (C-3'''), 122.3 (C-4'''), 125.5 (C-2'''/6'''), 128.7 (C-4'), 129.1 (C-3'''/5'''), 129.4 (C-3'/5'), 132.3 (C-2'/6'), 135.6 (C-1'''), 137.3 (C-1'), 137.4 (C-3), 136.79 (C-4), 147.7 (C-5), 149.6 (C-2''), 154.7 (C-5''), 169.4 (C-2). HRMS (ESI) Calcd for C₂₂H₁₅BrNaO₃ (M+Na), 431.0082; Found, 431.0092.

(Z)-3-Benzyl-5-((5-(2-nitrophenyl)furan-2-yl)methylene)furan-2(5H)-one (16a). Yield 48% (Z:E 63:37), diastereoisomers obtained after purification by column chromatography, eluent DCM /hexane (1:1 v/v). Dark yellow solid, mp 114.4-114.9 °C, isolated by HPLC, eluent DCM /hexane (3:1 v/v) and recrystallization with a mixture of DCM /hexane. IR (KBr) 3064, 3030, 2923, 1757, 1604, 1525, 1495, 1456, 1352, 1270, 1019 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 3.72 (s, 2H, H-7), 5.91 (s, 1H, H-6), 6.79 (d, 1H, H-4''; *J*_{4''/3''} 3.6 Hz), 6.93 (s, 1H, H-4), 7.08 (d, 1H, H-3''; *J*_{3''/4''} 3.6 Hz), 7.24-7.77 (m, 9H, H-2'''/6''', H-3'''/5''', H-2'/6', H-3'/5', H-4'). ¹³C NMR (75 MHz, CDCl₃), δ 31.8 (C-7), 100.2 (C-6), 113.0 (C-4''), 116.6 (C-3'''), 123.4 (C-3), 124.0 (C-6'''), 127.0 (C-4'), 128.7 (C-3'''), 128.8 (C-5'''), 128.9 (C-3'/5'), 129.0 (C-2'/6'), 132.9 (C-4'''), 133.2 (C-1'''), 137.0 (C-1'), 138.0 (C-4), 146.6 (C-5), 147.4 (C-2'''), 149.1 (C-5''), 150.2 (C-2''), 169.7 (C-2). HRMS (ESI) Calcd for C₂₂H₁₅NNaO₅ (M+Na), 396.0842; Found, 396.0850.

(E)-3-Benzyl-5-((5-(2-nitrophenyl)furan-2-yl)methylene)furan-2(5H)-one (16b). Yellow solid, mp 122.4-122.7 °C, purified by HPLC, eluent DCM /hexane (3:1 v/v) and recrystallization with a mixture of DCM /hexane. IR (KBr) 3063, 2925, 1763, 1605, 1529, 1461, 1359, 1283, 1030 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 3.78 (s, 2H, H-7), 6.29 (s, 1H, H-6), 6.56 (d, 1H, H-3''; *J*_{4''/3''} 3.3 Hz), 6.75 (d, 1H, H-4''; *J*_{4''/3''} 3.3 Hz), 7.25-7.70 (m, 10H, H-2'''/6''', H-3'''/5''', H-2'/6', H-3'/5', H-4', H-4). ¹³C NMR (75 MHz, CDCl₃), δ 32.3 (C-7), 101.8 (C-6), 112.4 (C-4''), 116.4 (C-3''), 123.6 (C-3), 124.2 (C-6'''), 127.0 (C-4'), 128.7 (C-3'''), 129.0 (C-5'''), 129.0 (C-3'/5'), 129.3 (C-2'/6'), 132.2 (C-4'''), 135.6 (C-1'''), 137.5 (C-1'), 146.6 (C-2'''), 146.8 (C-4), 147.7 (C-5''), 150.3 (C-5), 151.0 (C-2''), 169.5 (C-2). HRMS (ESI) Calcd for C₂₂H₁₅NNaO₅ (M+Na), 396.0842; Found, 396.0850.

(Z)-3-benzyl-4-isopropyl-5-(4-nitrobenzylidene)furan-2(5H)-one (17). Yield 36%, yellow solid, mp 90.0-91.0 °C. IR (neat) 3070, 2925, 2850, 1768, 1518, 1435, 1110, 1075 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 1.34 (d, 6H, H-2''; *J* 7.0 Hz), 3.19 (sept, 1 H, H-1''; *J* 7.0 Hz), 3.83 (s, 2H, H-7), 6.20 (s, 1H, H-6), 7.21-7.34 (m, 5H, H-2'/H-6'), 7.93 (d, 2H, H-2'''/H-6'''; *J* 9.1 Hz), 8.22 (d, 2H, H-3'''/H-5'''; *J* 9.1 Hz). ¹³C NMR (75 MHz, CDCl₃), δ 21.8 (C-2''), 26.8 (C-1''), 30.1 (C-7), 107.7 (C-6), 123.1 (C-3'''/C-5'''), 127.1 (C-4'), 128.6 (C-3'/C-5'), 129.1 (C-2'/C-6'), 131.1 (C-2'''/C-6'''), 135.0 (C-1'''), 137.8 (C-1'), 139.2 (C-3), 139.9 (C-4'''), 149.9 (C-4), 157.6 (C-5), 170.0 (C-2). HRMS (ESI) Calcd for C₂₁H₁₉NNaO₄⁺ (M+Na), 372.1206; Found, 372.1193.

5(Z)-3-benzyl-4-isopropyl-5-(3-nitrobenzylidene)furan-2(5H)-one (18). Yield 83%, yellow solid, mp 124.5-125.5 °C. IR (neat) 3100, 3080, 1766, 1645, 1615, 1531, 1495, 1350, 1099, 1033 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 1.34 (d, 6H, H-2''; *J* 7.0 Hz), 3.19 (sept, 1H, H-1''; *J* 7.0 Hz), 3.83 (s, 2H, H-7), 6.20 (s, 1H, H-6), 7.20-7.32 (m, 5H, H-2'-H-6'), 7.57 (t, 1H, H-5'''; *J* 8.0 Hz), 8.13 (ddd, 1H, H-4'''; *J* 2.1 Hz, *J* 0.9 Hz, *J* 8.0 Hz), 8.17 (br d, 1H, *J* 8.0 Hz, H-6'''), 8.44 (t, 1H, H-2'''; *J* 2.1 Hz). ¹³C NMR (75 MHz, CDCl₃), δ 21.8 (C-2''), 26.8 (C-1''), 29.9 (C-7), 107.7 (C-6), 123.1 (C-4'''), 125.1 (C-6'''), 126.9 (C-4'), 128.5 (C-3'/C-5'), 129.0 (C-2'/C-6'), 130.0 (C-3'''), 135.1 (C-1'''), 136.0 (C-3), 138.0 (C-2'''), 136.7 (C-1'), 139.1 (C-4), 157.6 (C-5), 149.2 (C-5'''), 170.0 (C-2). HRMS (ESI) Calcd for C₂₁H₁₉NNaO₄⁺ (M+Na), 372.1206; Found, 372.1192.

(Z)-3-Benzyl-4-isopropyl-5-(4-trifluoromethylbenzylidene)furan-2(5H)-one (19). Yield 22%, White solid, mp 83.5-84.0 °C. IR (neat) 3030, 2975, 2930, 1769, 1649, 1615, 1495, 1416, 1324, 1165 cm⁻¹. ¹H NMR (300 MHz, CDCl₃), δ 1.34 (d, 6H, H-2''; *J* 7.2 Hz), 3.18 (sept, 1H, H-1''; *J* 7.2 Hz), 3.83 (s, 2H, H-7), 6.21 (s, 1H, H-6), 7.20-7.32 (m, 5H, H-2'-H-6'), 7.62 (d, 2H, H-3'''/H-5'''; *J* 8.4 Hz), 7.88 (d, 2H, H-2'''/H-6'''; *J* 8.4 Hz). ¹³C NMR (75 MHz, CDCl₃), δ 21.9 (C-2''), 26.8 (C-1''), 29.9 (C-7), 108.7 (C-6), 125.9 (q, C-3'''/C-5'''; *J*_{C,F} 3.8 Hz), 126.0 (CF₃), 126.9 (C4'), 128.5 (C-2'/C-6'), 128.9 (C-3'/C-5'), 130.5 (C-4'''), 130.7 (C-2'''/C-6'''), 134.0 (C-3), 136.5 (C-1'''), 136.8 (C-1'), 138.0 (C-4), 157.8 (C-5), 170.3 (C-2). HRMS (ESI) Calcd for C₂₂H₁₉F₃O₂ (M+1), 373.1371. HRMS (ESI) Calcd for C₂₂H₁₉F₃NaO₂⁺ (M+Na), 395.1229; Found, 395.1220.

(Z)-3,4-Dichloro-5-(3,4-metylenedioxybenzylidene)furan-2(5H)-one (20). Yield 70%, yellow solid, mp 159.5-160.2 °C. IR (neat) 3078, 2919, 1769, 1652, 1618, 1570, 1503, 1453, 1269,

1254, 1108 cm^{-1} . ^1H NMR (300 MHz, CDCl_3), δ 6.04 (s, 2H, $-\text{OCH}_2\text{O}-$), 6.32 (s, 1H, H-6), 6.84 (d, 1H, H-5'; J 8.1 Hz), 7.18 (dd, 1H, H-6'; J 8.1 Hz, J 1.8 Hz), 7.44 (d, 1H, H-2'; J 1.8 Hz). ^{13}C NMR (75 MHz, CDCl_3), δ 102.1 ($-\text{OCH}_2\text{O}-$), 109.2 (C-5'), 110.5 (C-2'), 113.0 (C-6), 126.2 (C-6'), 127.4 (C-3), 130.2 (C-1'), 132.3 (C-4), 148.7 (C-3'), 149.8 (C-4'), 161.5 (C-5), 162.4 (C-2). HRMS (ESI) Calcd for $\text{C}_{12}\text{H}_5\text{Cl}_2\text{O}_4^- (\text{M}-\text{H}^+)$, 282.9570; Found 282.9574.

(Z)-5-(Benzo[d][1,3]dioxol-5-ylmethylene)-3,4-dibromofuran-2(5H)-one (21). Yield 65%, yellow solid, mp 170.7-171.5 $^\circ\text{C}$; IR (neat) 3078, 2918, 1769, 1650, 1569, 1502, 1269, 1254, 1106 cm^{-1} . ^1H NMR (300 MHz, CDCl_3), δ 6.05 (s, 2H, $-\text{OCH}_2\text{O}-$), 6.36 (s, 1H, H-6), 6.85 (d, 1H, H-5'; J 8.1 Hz), 7.22 (dd, 1H, H-6'; J 8.1 Hz, J 1.8 Hz), 7.46 (d, 1H, H-2'; J 1.8 Hz). ^{13}C NMR (75 MHz, CDCl_3), δ 101.8 ($-\text{OCH}_2\text{O}-$), 108.5 (C-5'), 109.0 (C-2'), 110.3 (C-6), 125.4 (C-6'), 127.5 (C-3), 129.9 (C-1'), 132.0 (C-4), 148.7 (C-3'), 149.8 (C-4'), 161.1 (C-2), 162.0 (C-5). HRMS (ESI) Calcd for $\text{C}_{12}\text{H}_5\text{Br}_2\text{NaO}_4^- (\text{M}-\text{H}^+)$, 370.8560; Found 370.8562

(Z)-3,4-Dibromo-5-(4-nitrobenzylidene)furan-2(5H)-one (22). Yield 13%, yellow solid, mp 188.1-189.0 $^\circ\text{C}$. IR (neat) 3040, 1725, 1593, 1609, 1518, 1345, 1108, 1024 cm^{-1} . ^1H NMR (300 MHz, CDCl_3), δ 6.77 (s, 1H, H-6), 8.01 (d, 2H, H-3'/H-5'; J 9.0 Hz), 8.28 (d, 2H, H-2'/H-6'; J 9.0 Hz). ^{13}C NMR (75 MHz, CDCl_3), δ 105.0 (C-6), 119.1 (C-3), 123.8 (C-3'/C-5'), 129.8 (C-2'/C-6'), 135.3 (C-4), 137.0 (C-1'), 146.0 (C-5), 148.5 (C-4'), 158.3 (C-2). HRMS (ESI) Calcd for $\text{C}_{11}\text{H}_4\text{Br}_2\text{NO}_4^- (\text{M}-\text{H}^+)$, 371.8513; Found 371.8510.

(Z)-3-Benzyl-5-((E)-3-phenylallylidene)furan-2(5H)-one (23). Yield 32%, yellow solid, mp 112.0 – 112.8 $^\circ\text{C}$. IR (neat) 3086, 3059, 3030, 2923, 1759, 1630, 1594, 1570, 1496, 1457, 1328, 1032 cm^{-1} . ^1H NMR (300 MHz, CDCl_3), δ 3.73 (br s, 2H, H-9), 5.85 (d, 1H, H-6; $J_{6/7}$ 11.4 Hz), 6.74 (d, 1H, H-8; $J_{8/7}$ 15.6 Hz), 6.89 (t, 1H, H-4; $J_{4/9}$ 1.3 Hz), 7.25-7.40 (m, 9H, H-7, H-2'/H-6', H-3'/H-5'', H-3''/H-5'), 7.42-7.46 (m, 2H, H-2''/6''). ^{13}C NMR (75 MHz, CDCl_3), δ 32.2 (C-9), 113.9 (C-6), 121.9 (C-7), 127.2 (C-3), 127.3 (C-4'), 127.4 (C-2''/6''), 129.0 (C-3''/5''), 129.1 (C-2'/6'), 129.2 (C-3'/5'), 129.3 (C-4''), 133.47 (C-1''), 136.73 (C-1'), 137.45 (C-8), 137.71 (C-4), 148.19 (C-5), 169.98 (C-2). HRMS (ESI) Calcd $\text{C}_{20}\text{H}_{16}\text{NaO}_2 (\text{M}+\text{Na})$, 311.1043; Found 311.1033

Biological assays. The ability to interfere with the photosynthetic electron transport chain was evaluated on photosynthetically active thylakoid membranes isolated from market spinach (*Spinacia oleracea* L.) leaves. Deveined plant material was resuspended in 5 mL g^{-1} of ice-cold 20 mM N-tris(hydroxymethyl)methylglycine (Tricine)-NaOH buffer (pH 8.0) containing 10 mM NaCl, 5 mM MgCl_2 , and 0.4 M sucrose, and homogenized for 30 s in a blender at maximal speed. The homogenate was filtered through surgical gauze, and the filtrate was centrifuged at 4 $^\circ\text{C}$ for 1 min at 500 g; the supernatant was further centrifuged for 10 min at 1500 g. Pelleted chloroplasts were osmotically swollen by resuspension in sucrose-lacking buffer. The suspension was immediately diluted 1:1 with sucrose-containing buffer, kept on ice in the dark, and used within a few hours from preparation. Following proper dilution with 80% (v/v) acetone, the absorbance of each sample was determined at 645 and 663 nm, and the chlorophyll content was calculated on the basis of Arnon's formula.

The basal rate of photosynthetic electron transport was measured following light-driven ferricyanide reduction. Aliquots of membrane preparations corresponding to 15 μg of chlorophyll were incubated at 24 $^{\circ}\text{C}$ in 1 mL cuvettes containing 20 mM Tricine-NaOH buffer (pH 8.0), 10 mM NaCl, 5 mM MgCl_2 , 0.2 M sucrose, and 1 mM $\text{K}_3\text{Fe}(\text{CN})_6$. The assay was initiated by exposure to saturating light ($800 \mu\text{mol m}^{-2} \text{s}^{-1}$), and the rate of ferricyanide reduction was measured at 1 min intervals for 10 min against an exact blank at 420 nm. Activity was calculated over the linear portion of the curve from a molar extinction coefficient of $1000 \text{ M}^{-1} \text{ cm}^{-1}$. The mean value (\pm SD) for untreated controls was $73.2 \pm 7.4 \text{ nmol s}^{-1} (\text{mg of chlorophyll})^{-1}$ in 18 independent thylakoid membrane preparations. Nostoclid analogues were dissolved in DMSO so as to obtain 10 mM solutions that were then water diluted, as appropriate. Their effect upon the photosynthetic electron transport was evaluated in parallel assays in which the compounds were added to the reaction mixture to a final concentration of 1, 2, 5, 10, 20, 50 or 100 μM . For each dose the assay was carried out in triplicate. The concentrations causing 50% inhibition (IC_{50}) were estimated utilizing the linear regression equation of activity values, expressed as percentage of untreated controls, plotted against the logarithm of inhibitor concentration. Confidence limits were computed according to Snedecor and Cochran.²⁶ Phosphorylating electron flow was assessed under the same conditions, but in the presence of 0.5 mM ADP and 2 mM K_2HPO_4 . Unless otherwise indicated, uncoupled activity was measured following the addition of 0.5 mM NH_4Cl to the basal reaction mixture containing aliquots of membrane preparations corresponding to 7.5 μg chlorophyll.

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